

Announcements

Wednesday, November 25, 2009

MasteringChemistry due dates (all at 11:59 pm)

- Ch 7: Wed, Nov 25
- Ch 8: Wed, Dec 2
- Ch 9: Fri, Dec 4

Ionization energy

Ionization energy (IE): energy required to remove 1 electron from an atom or ion



Trends in first ionization energy (IE₁):

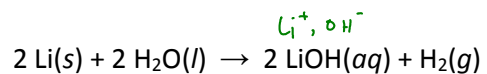


(additional protons across a period will increase attraction to the electrons)



(an additional shell shields the valence electrons from the nuclear attraction)

Ionization energy and Electron affinity



oxidations



K is the strongest reducing agent of these

Reactions of the Alkali Metals with Water



Lithium



Sodium



Potassium

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Electron affinity (EA): energy change associated with an electron being added to a neutral atom.

Li EA = -60 kJ/mol

O EA = -141 kJ/mol

F EA = -328 kJ/mol