

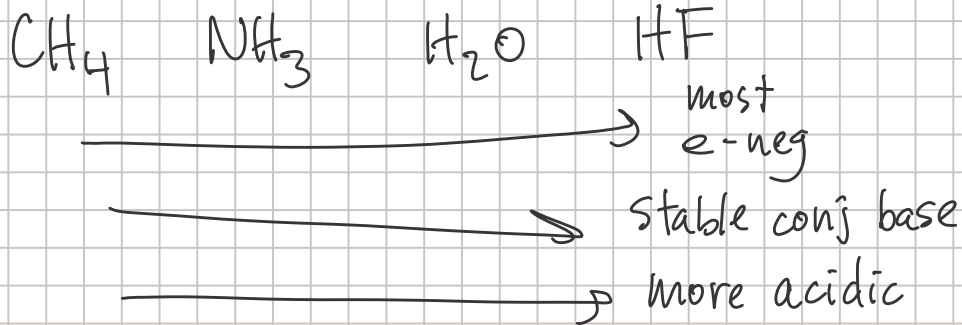
# Ch 1

Note Title

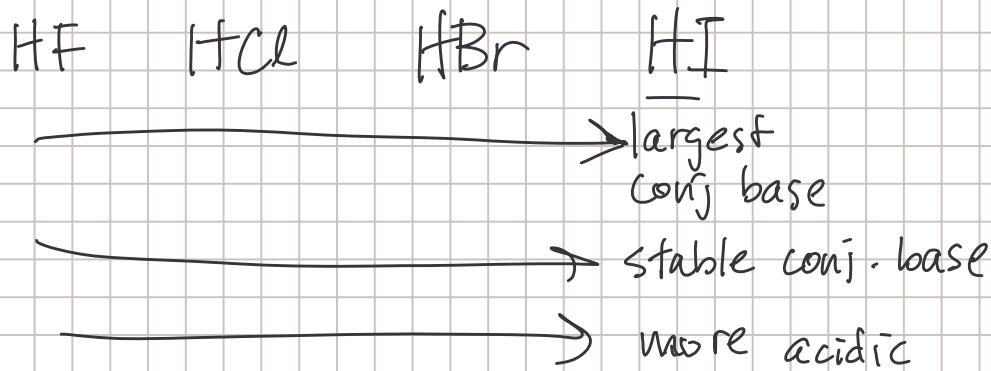
8/30/2005

## Structural effects on acidity

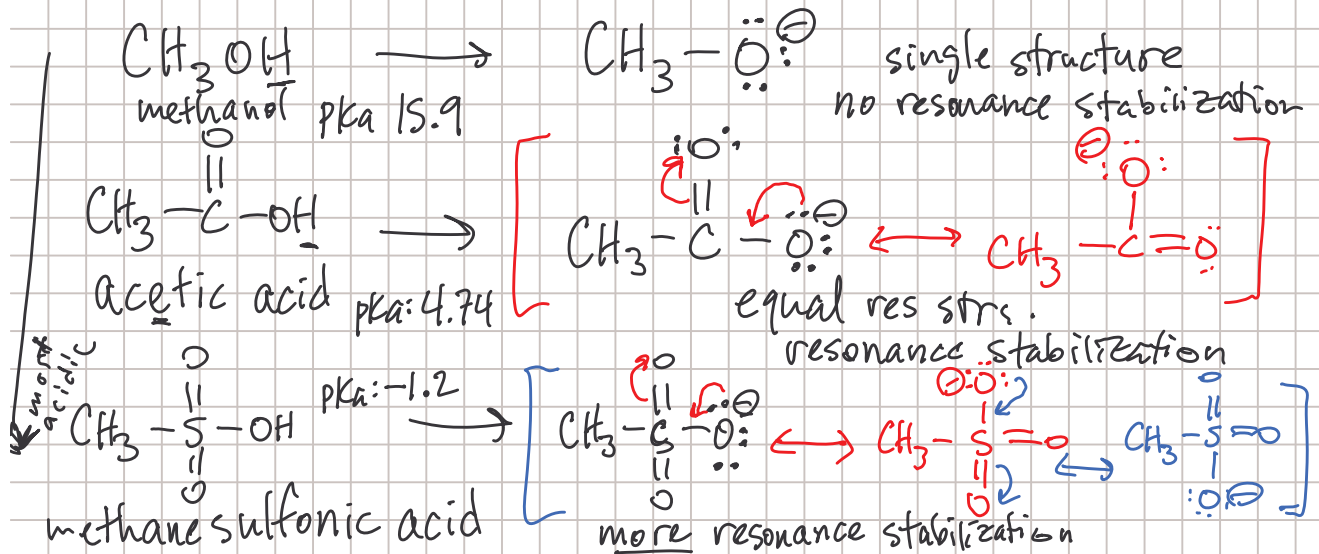
### Electronegativity



### Size



### Resonance stabilization

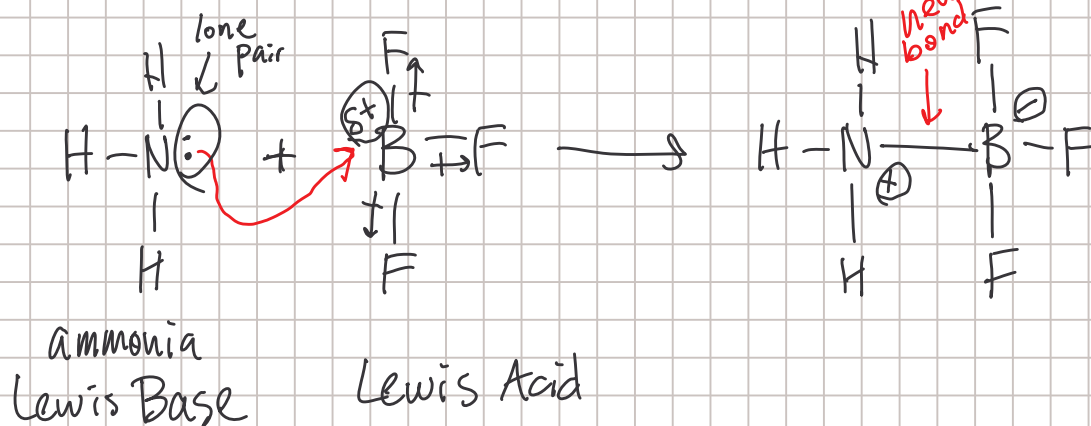


Smaller  $pK_a$  = stronger acid  
 larger  $K_a$  = stronger acid

Lewis acids/bases      no proton xfer necessary

L. acid : accepts electrons (electrophile)

L. base : donates electrons (nucleophile)



nucleophile  $\delta^-$   
 electrophile  $\delta^+$

