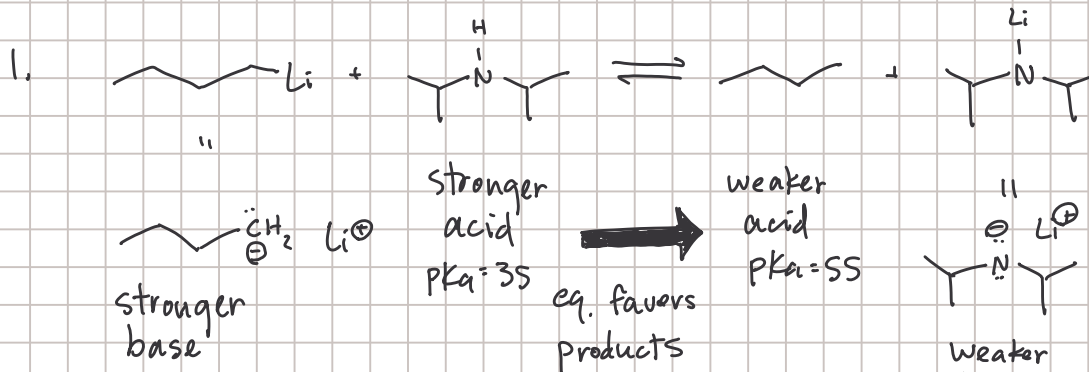


Ch 15 group work.

Note Title

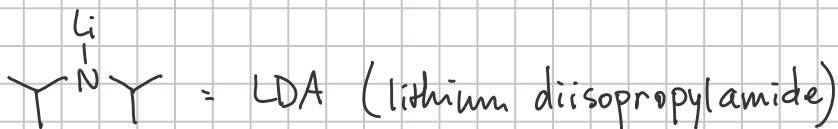
2/1/2006



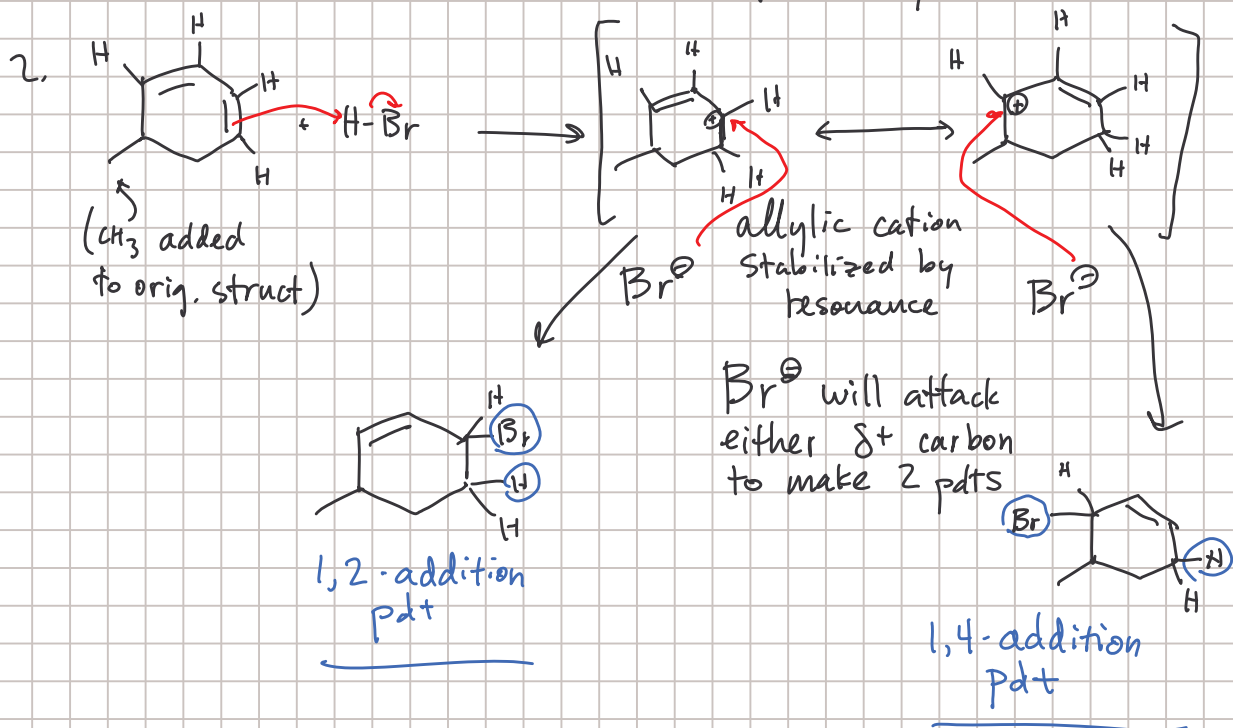
(st. acid more likely to dissociate.
st. base more likely to become protonated)

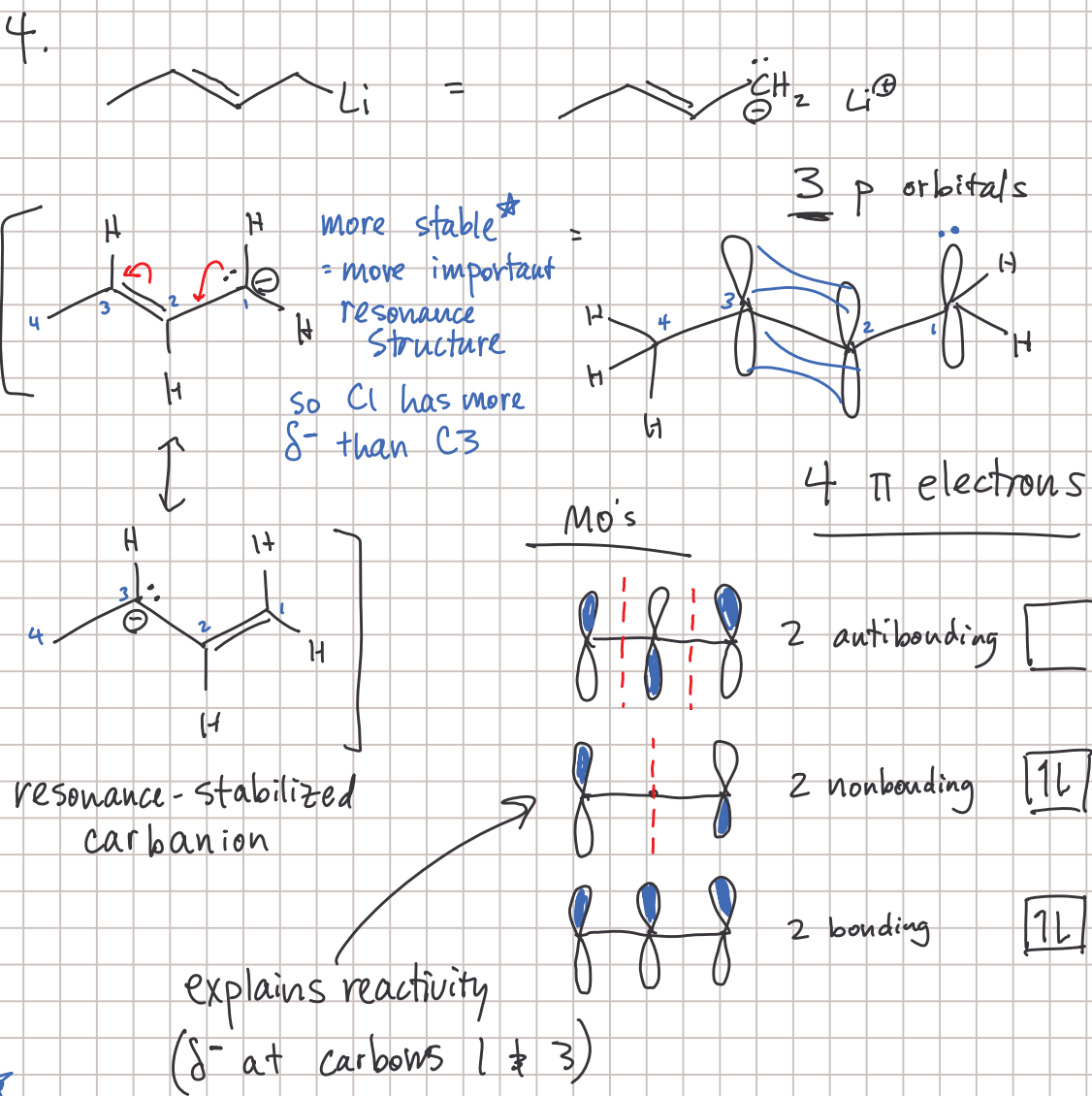
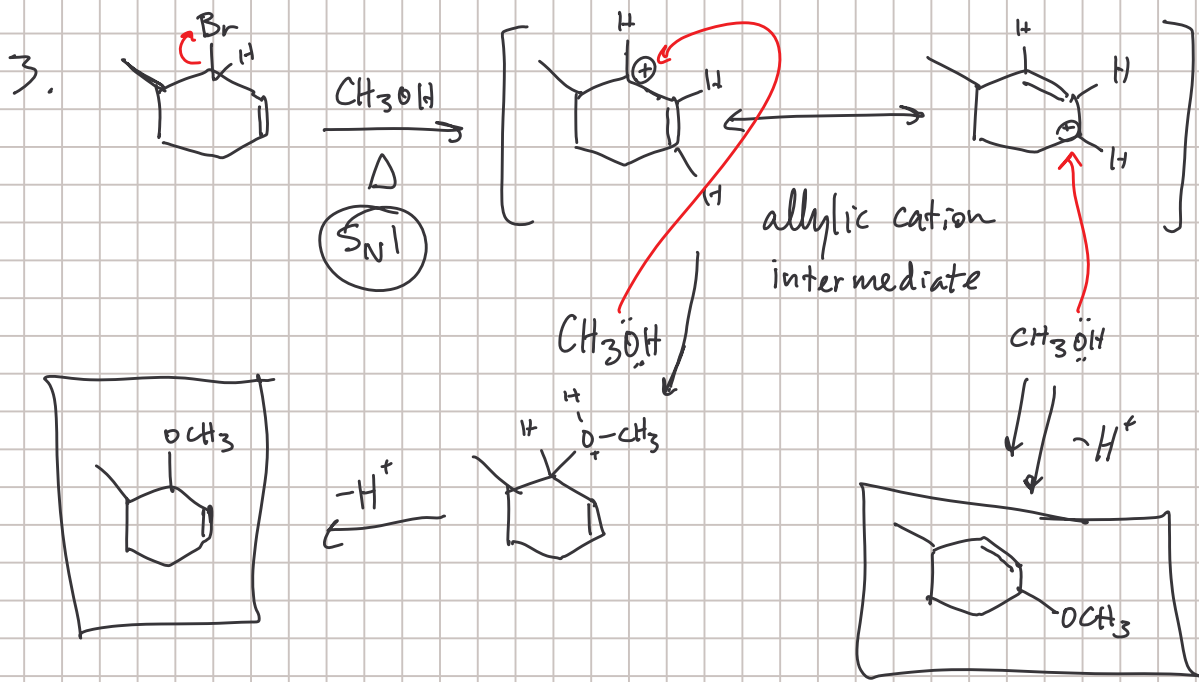
Remember, a

C-Li bond is always ionic - exactly the same as $C^- Li^+$
(only drawn as C-Li for convenience)



a strong, bulky base (won't substitute)
- not a nucleophile





* \ominus charges are destabilized by extra e^- donation from neighboring carbons, so the res. str. w/ the $1^\circ \ominus$ is more stable