

**Nomenclature Worksheet**  
**Chem 1020**

Name \_\_\_\_\_

1. Write formulas for the following compounds:

- |   |   |
|---|---|
| a) LiBr   | b) $\text{CuCl}_2$  |
| c) $\text{Ba}_3(\text{PO}_4)_2$                             | d) $\text{Cs}_2\text{CO}_3$   |
| e) $\text{Mg}(\text{NO}_3)_2$                               | f) $\text{K}_2\text{SO}_4$  |
| g) $\text{Fe}_2\text{O}_3$                                  | h) $\text{AlI}_3$   |
| i) $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$              | j) $\text{Ca}(\text{OH})_2$   |
| k) ZnO (Zn is always 2+ so it doesn't need a roman numeral) | l) $\text{Ag}_2\text{S}$ (Ag is always 1+ so it doesn't need a roman numeral) |
| m) $\text{CCl}_4$   | n) $\text{SF}_6$  |
| o) $\text{N}_2\text{O}_5$                                   | p) $\text{PBr}_3$   |
| q) $\text{NaHCO}_3$   | r) $\text{HNO}_3$   |
| s) HF (aq)  |   |

2. Name the following compounds with the *correct spelling*:

- |                           |   |
|---------------------------|---|
| a) ammonium sulfate       | b) iron (III) sulfide                                   |
| c) calcium acetate        | d) sodium phosphate                                     |
| e) lithium sulfate        | f) zinc nitrate (Zn is always 2+)                       |
| g) aluminum carbonate     | h) copper (I) phosphide                                 |
| i) potassium hydroxide    | j) calcium bicarbonate<br>or calcium hydrogen carbonate |
| k) diphosphorus pentoxide | l) strontium phosphate                                  |
| m) carbon monoxide        | n) dichlorine monoxide                                  |
| o) magnesium carbonate    | p) water or dihydrogen monoxide                         |
| q) phosphoric acid        | r) hydroiodic acid                                      |