

Reactions

Name(s):

Date:

Lab Section:

DATA

1. Enter your data for the CaCl_2 reactions from Procedure 1 in the table below.

Test Tube	1	2	3	4	5	6	7
	MgSO_4	$(\text{NH}_4)_2\text{C}_2\text{O}_4$	KNO_3	Na_3PO_4	KBr	NaOH	K_2CO_3
Observations							

2. Record your observations for the gas-evolving reactions from Procedure 2.
3. Record your observations for the acid-base reactions from Procedure 3. How many drops of sodium hydroxide were needed for the permanent color change for each acid?
4. Record your observations for the redox reactions from Procedure 4.

QUESTIONS

1. For **all of the reactions** observed in each procedure, write the balanced molecular and net ionic equations. Make sure to include phase labels after each reactant and product.

Procedure 1

Procedure 2

Procedure 3

Procedure 4

2. For each reaction observed in Procedure 4, write out the balanced net ionic equation with the oxidation numbers of each species above it. Identify the oxidizing agent and reducing agent of each reaction.